**Conjugate acid/base pairs (Level 2) examiners tips: Read these please!**

• an increase in conductivity is due to an increase in the concentration of ions

• a weak acid only partially dissociates

• a low pH indicates that the substance is an acid with a high concentration of H3O+ ions

• a high pH indicates that the substance is a base with a high concentration of OH- ions

**Also…don’t be daft!  
Amphiprotic:** a substance which can both accept and donate protons (H+ ions)

*eg water acting as a base*

HCl + H2O 🡪 Cl- + H3O+

*eg water acting as an acid*

NH3 + H2O ⇌ NH4+ + OH-

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